

S1 Periodic Table

- Each lesson of this topic you will need:
 - Show me board and pen (collect on your way in)
 - iPad with Showbie opened to the S1 Chemistry Tab in Showbie

Daily Review

What do you already know?

1.   1p2 Science
2.  S1 Chemistry
3.  Lesson 1
4. S1 Chem Lesson 1 Daily Review



Periodic Table

S1 SCIENCE

1	2											18	19																																																																																						
3	4											5	6	7	8	9	10	17	16																																																																																
11	12											13	14	15	14	15	16	17	18																																																																																
19	20											21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36																																																																								
37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54	55	56	57	58	59	60	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86	87	88	89	90	91	92	93	94	95	96	97	98	99	100																																				
101	102	103	104	105	106	107	108	109	110	111	112	113	114	115	116	117	118	119	120	121	122	123	124	125	126	127	128	129	130	131	132	133	134	135	136	137	138	139	140	141	142	143	144	145	146	147	148	149	150	151	152	153	154	155	156	157	158	159	160	161	162	163	164	165	166	167	168	169	170	171	172	173	174	175	176	177	178	179	180	181	182	183	184	185	186	187	188	189	190	191	192	193	194	195	196	197	198	199	200

Daily Review

What do you already know?

1. 🧑‍🔬 S1 Chemistry
2. 🧱 Lesson 2
3. S1 Chem Lesson 2
Daily Review

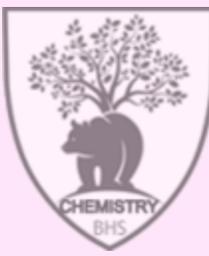


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We are learning about

Atoms and Molecules



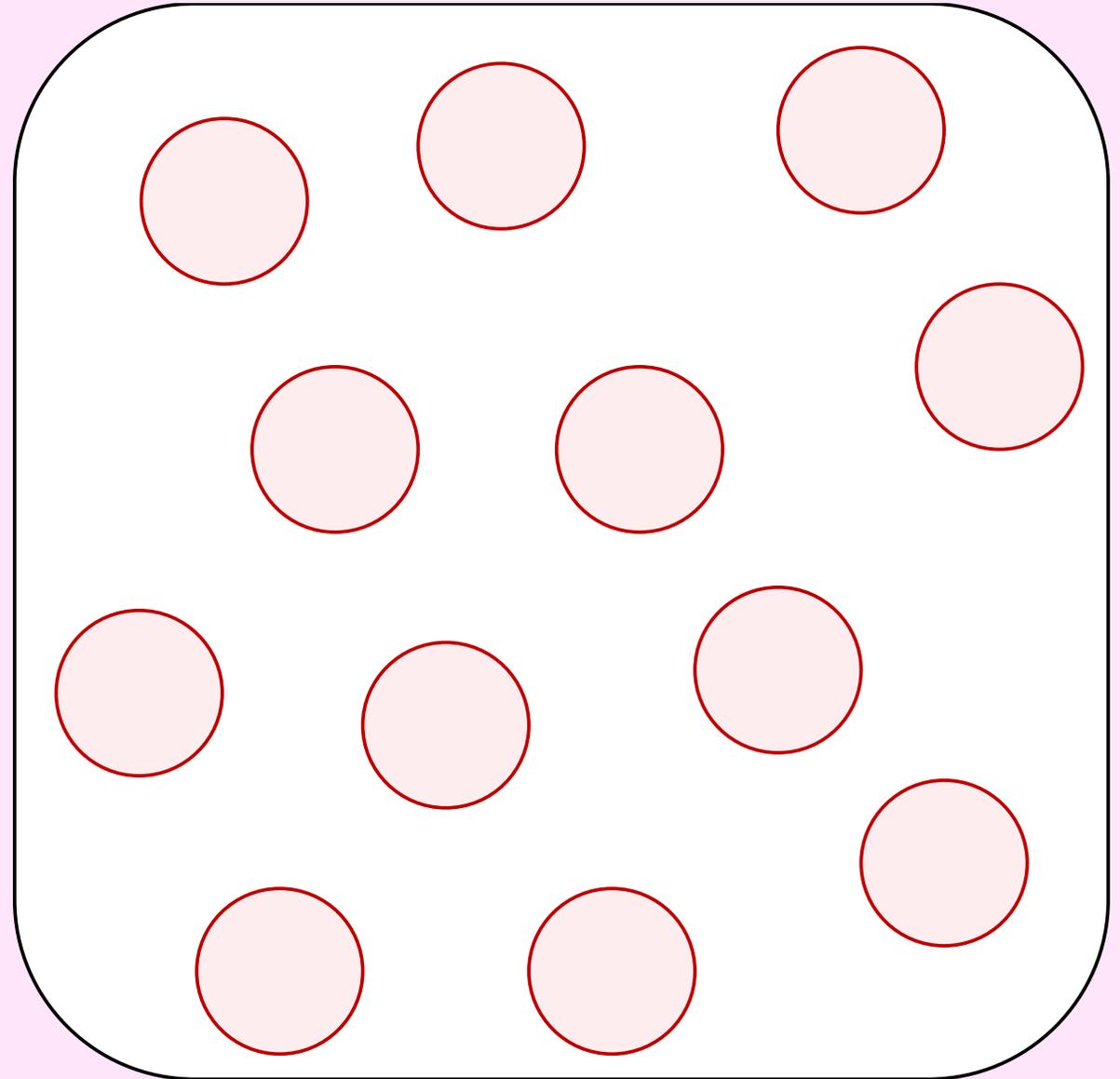
- State what is meant by an atom and a molecule.
- Describe what was special about how Dmitri Mendeleev created the first periodic table
- Explain what an element is and where they are located.
- State where on the Periodic Table you find metals and non-metals.

What is an atom?

Atoms

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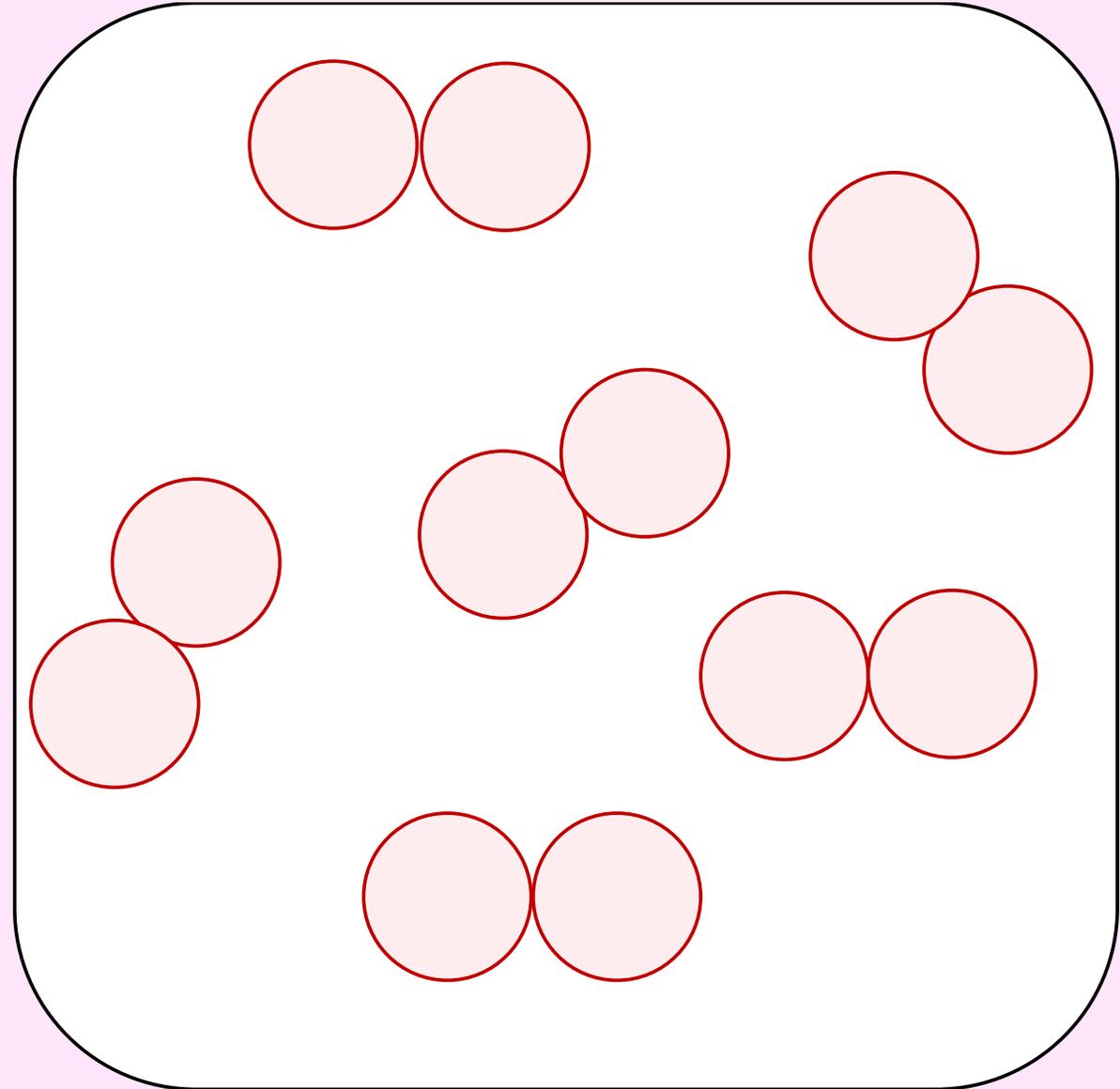
Everything in the world is made up from tiny particles called **atoms**.



Molecule

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A molecule is when
two or more atoms
are bonded together.



Tasks - Lesson 1

1. Attempt **tasks 2, 3, 4, and 6** on the Showbie document
2.  **Chemistry Lesson 1**
Plenary review



Daily Review

What do you already know?

1. 🧑‍🎓 S1 Chemistry
2. 🧱 Lesson 2
3. S1 Chem Lesson 2
Daily Review

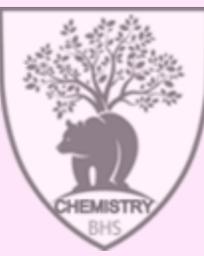


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We are learning about

Elements



I can describe and compare the key words:

Atom

Molecule

Element

Periodic Table

Elements

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Elements are made of one type of atom and cannot be broken down.

All elements are shown on the periodic table.

Dmitri Mendeleev created the first periodic table, and predicted the size and properties of elements not yet discovered

Elements

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Elements	Not Elements
Oxygen Copper Gold Carbon Iron	Water Carbon dioxide Salt Air Sand

Oxygen

Water

Copper

Carbon dioxide

Gold

Carbon

Air

Salt

Iron

Sand

Tasks - Lesson 2

1. Attempt tasks on lesson 2 worksheet
2.  Chemistry Lesson 2 Plenary review



THE GENIUS OF
MENDELEEV'S
PERIODIC TABLE



Task - Pairs

Using the cards, arrange, and predict the number and colors of the missing blanks.

Daily Review

1. What is an atom?
2. What is a molecule?
3. What is the periodic table for?
4. If 10 students in a class of 20 have blue eyes, what percentage of students have blue eyes?

Daily Review

1. What is an atom?
2. What is a molecule?
3. Who invented the periodic table?
4. 5 students in a class of 20 have brown eyes, what percentage of students have brown eyes?

Periodic Table of the Elements

1 IA 1A																	18 VIIIA 8A
1 H Hydrogen 1.008	2 IIA 2A											13 IIIA 3A	14 IVA 4A	15 VA 5A	16 VIA 6A	17 VIIA 7A	2 He Helium 4.003
3 Li Lithium 6.941	4 Be Beryllium 9.012											5 B Boron 10.811	6 C Carbon 12.011	7 N Nitrogen 14.007	8 O Oxygen 15.999	9 F Fluorine 18.998	10 Ne Neon 20.180
11 Na Sodium 22.990	12 Mg Magnesium 24.305	3 IIIB 3B	4 IVB 4B	5 VB 5B	6 VIB 6B	7 VIIB 7B	8 VIII 8	9 VIII 8	10 VIII 8	11 IB 1B	12 IIB 2B	13 Al Aluminum 26.982	14 Si Silicon 28.086	15 P Phosphorus 30.974	16 S Sulfur 32.066	17 Cl Chlorine 35.453	18 Ar Argon 39.948
19 K Potassium 39.098	20 Ca Calcium 40.078	21 Sc Scandium 44.956	22 Ti Titanium 47.867	23 V Vanadium 50.942	24 Cr Chromium 51.996	25 Mn Manganese 54.938	26 Fe Iron 55.845	27 Co Cobalt 58.933	28 Ni Nickel 58.693	29 Cu Copper 63.546	30 Zn Zinc 65.38	31 Ga Gallium 69.723	32 Ge Germanium 72.631	33 As Arsenic 74.922	34 Se Selenium 78.971	35 Br Bromine 79.904	36 Kr Krypton 83.798
37 Rb Rubidium 85.468	38 Sr Strontium 87.62	39 Y Yttrium 88.906	40 Zr Zirconium 91.224	41 Nb Niobium 92.906	42 Mo Molybdenum 95.95	43 Tc Technetium 98.907	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.906	46 Pd Palladium 106.42	47 Ag Silver 107.868	48 Cd Cadmium 112.414	49 In Indium 114.818	50 Sn Tin 118.711	51 Sb Antimony 121.760	52 Te Tellurium 127.6	53 I Iodine 126.904	54 Xe Xenon 131.294
55 Cs Cesium 132.905	56 Ba Barium 137.328	57-71	72 Hf Hafnium 178.49	73 Ta Tantalum 180.948	74 W Tungsten 183.84	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.217	78 Pt Platinum 195.085	79 Au Gold 196.967	80 Hg Mercury 200.592	81 Tl Thallium 204.383	82 Pb Lead 207.2	83 Bi Bismuth 208.980	84 Po Polonium [208.982]	85 At Astatine 209.987	86 Rn Radon 222.018
87 Fr Francium 223.020	88 Ra Radium 226.025	89-103	104 Rf Rutherfordium [261]	105 Db Dubnium [262]	106 Sg Seaborgium [266]	107 Bh Bohrium [264]	108 Hs Hassium [269]	109 Mt Meitnerium [278]	110 Ds Darmstadtium [281]	111 Rg Roentgenium [280]	112 Cn Copernicium [285]	113 Nh Nihonium [286]	114 Fl Flerovium [289]	115 Mc Moscovium [289]	116 Lv Livermorium [293]	117 Ts Tennessine [294]	118 Og Oganesson [294]

Lanthanide Series

57 La Lanthanum 138.905	58 Ce Cerium 140.116	59 Pr Praseodymium 140.908	60 Nd Neodymium 144.243	61 Pm Promethium 144.913	62 Sm Samarium 150.36	63 Eu Europium 151.964	64 Gd Gadolinium 157.25	65 Tb Terbium 158.925	66 Dy Dysprosium 162.500	67 Ho Holmium 164.930	68 Er Erbium 167.259	69 Tm Thulium 168.934	70 Yb Ytterbium 173.055	71 Lu Lutetium 174.967
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Actinide Series

89 Ac Actinium 227.028	90 Th Thorium 232.038	91 Pa Protactinium 231.036	92 U Uranium 238.029	93 Np Neptunium 237.048	94 Pu Plutonium 244.064	95 Am Americium 243.061	96 Cm Curium 247.070	97 Bk Berkelium 247.070	98 Cf Californium 251.080	99 Es Einsteinium [254]	100 Fm Fermium 257.095	101 Md Mendelevium 258.1	102 No Nobelium 259.101	103 Lr Lawrencium [262]
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Element Cubes:

- What the element looks like
- Where it comes from
- What we use it for (this could be more than one side)
- What it does in the body
- How much of it there is in the world

There are many websites you can look at including <https://www.rsc.org/periodic-table>.

Daily Review

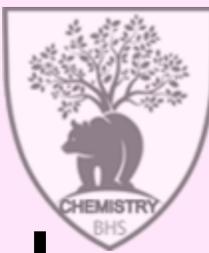
1. 5 students in a class of 20 have green eyes, what percentage of students have green eyes?
2. Use your periodic table to name: (P4)
 - a) Cu
 - b) K
 - c) Li
 - d) Na
 - e) Ca
 - f) Sr

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We are learning about

Metals and Non Metals



- I can describe the properties of metals
- I can describe the properties of non-metals
- I can use the periodic table to state if something is a metal or a nonmetal.
- I can explain how to test which metal an element is

Metals										Non-Metals													
1 H Hydrogen 1.008																	2 He Helium 4.003						
3 Li Lithium 6.941	4 Be Beryllium 9.012																	5 B Boron 10.811	6 C Carbon 12.011	7 N Nitrogen 14.007	8 O Oxygen 15.999	9 F Fluorine 18.998	10 Ne Neon 20.180
11 Na Sodium 22.990	12 Mg Magnesium 24.305																	13 Al Aluminum 26.982	14 Si Silicon 28.086	15 P Phosphorus 30.974	16 S Sulfur 32.066	17 Cl Chlorine 35.453	18 Ar Argon 39.948
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Metals are on the left of the zig-zag

Non-metals are on the right (except hydrogen).

1 H Hydrogen 1.008																	2 He Helium 4.003
3 Li Lithium 6.941	4 Be Beryllium 9.012											5 B Boron 10.811	6 C Carbon 12.011	7 N Nitrogen 14.007	8 O Oxygen 15.999	9 F Fluorine 18.998	10 Ne Neon 20.180
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Testing Metals

- Bunsen Burner
- Metals to test

Results: L3 Metals in Showbie
Chemistry, Periodic
table

Daily Review

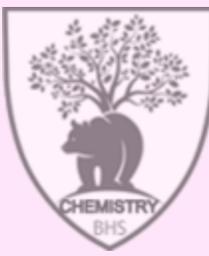
1. 4 students in a class of 25 have green eyes, what percentage of students have green eyes?
2. What color would the following metals burn? (P6)
 - a) Cu
 - b) K
 - c) Li
 - d) Na
 - e) Ca
 - f) Sr

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We are learning about the

Groups on the Periodic Table



- I can state the location of **alkali metals, transition metals, halogens and noble gases** on the Periodic Table.
- I can explain what happens to the **reactivity of elements** as you go down a group.

Elements are divided into **columns called groups.**

Elements in the **same group** have very **similar chemical properties.**

1 H Hydrogen 1.008	Alkali Metals										Transition Metals										2 He Helium 4.003	
3 Li Lithium 6.941	4 Be Beryllium 9.012	Halogens										Noble Gases										
11 Na Sodium 22.990	12 Mg Magnesium 24.305												13 Al Aluminum 26.982	14 Si Silicon 28.086	15 P Phosphorus 30.974	16 S Sulfur 32.066	17 Cl Chlorine 35.453	18 Ar Argon 39.948				
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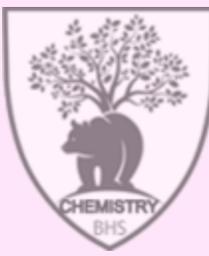
The further **down** the alkali metals group the element is, the **more reactive** it is with water.



Daily Review

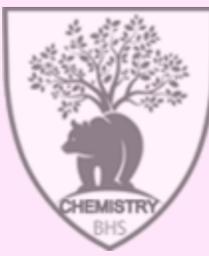
1. What are the 4 Groups on the periodic table?
2. What happens with alkali metals as we go further down the periodic table?
3. For each element, write down the name and group
 - a) Ir
 - b) Kr
 - c) Br
 - d) K

25/01/2026



We are learning about the

Groups on the Periodic Table



- I can state the location of **alkali metals, transition metals, halogens and noble gases** on the Periodic Table.
- I can describe properties of **halogens, transition metals and noble gases**

Nobel Gases

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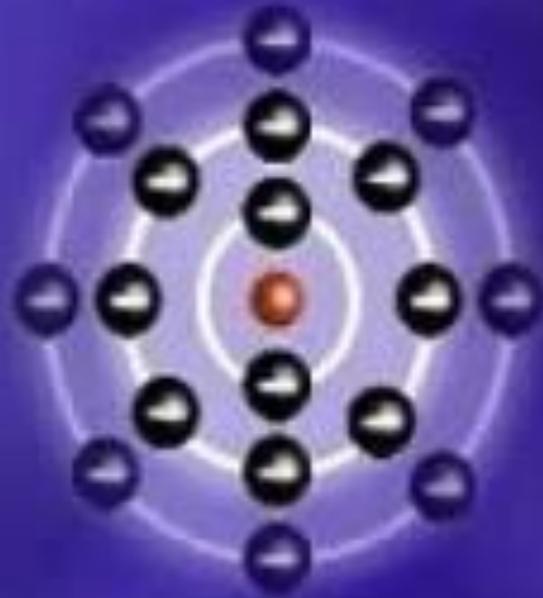
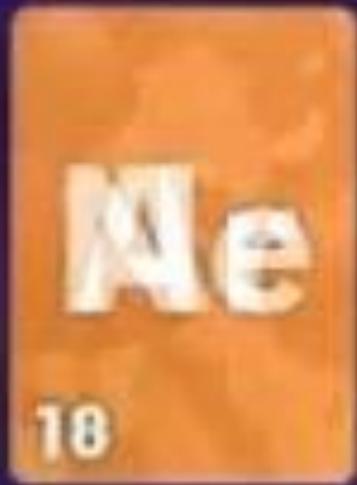
What are the 6 Nobel gasses?

What does helium do?

What does Neon do?

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Nobel Gases



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Halogens



Transition metals are the 'normal' metals.

Whiteboards - describe what metal is like

Property

Metals

Non-metals

Appearance

shiny

dull

**Electrical
Conductivity**

good

poor

**Heat
Conductivity**

good

poor

Malleability

malleable

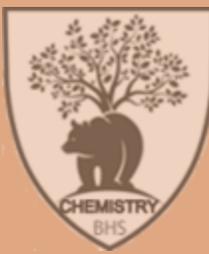
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1. Task in the Review
Tab on Teams

2. Top QR Code Quiz

3. Bottom QR Code
Questions (Jotter)

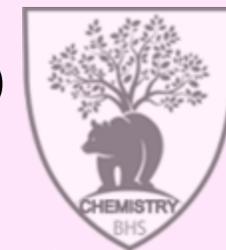




- State what is meant by an atom and a molecule.
- Describe what was special about how Dmitri Mendeleev created the first periodic table
- Explain what an element is and where they are located.
- Describe the properties of metals
- Describe the properties of non-metals
- I can use the periodic table to state if something is a metal or a nonmetal.
- Explain how to test which metal an element is
- state the location and properties of **alkali metals, transition metals, halogens and noble gases** on the Periodic Table.

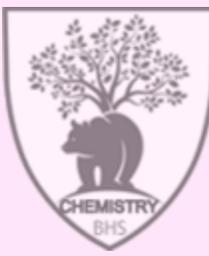
Daily Review

1. What are the 4 Groups on the periodic table?
2. What happens with Nobel gases as we go further down the periodic table?
3. For each element, write down the name, atomic number and group
 - a) Ir
 - b) Kr
 - c) Br
 - d) Ba
 - e) K



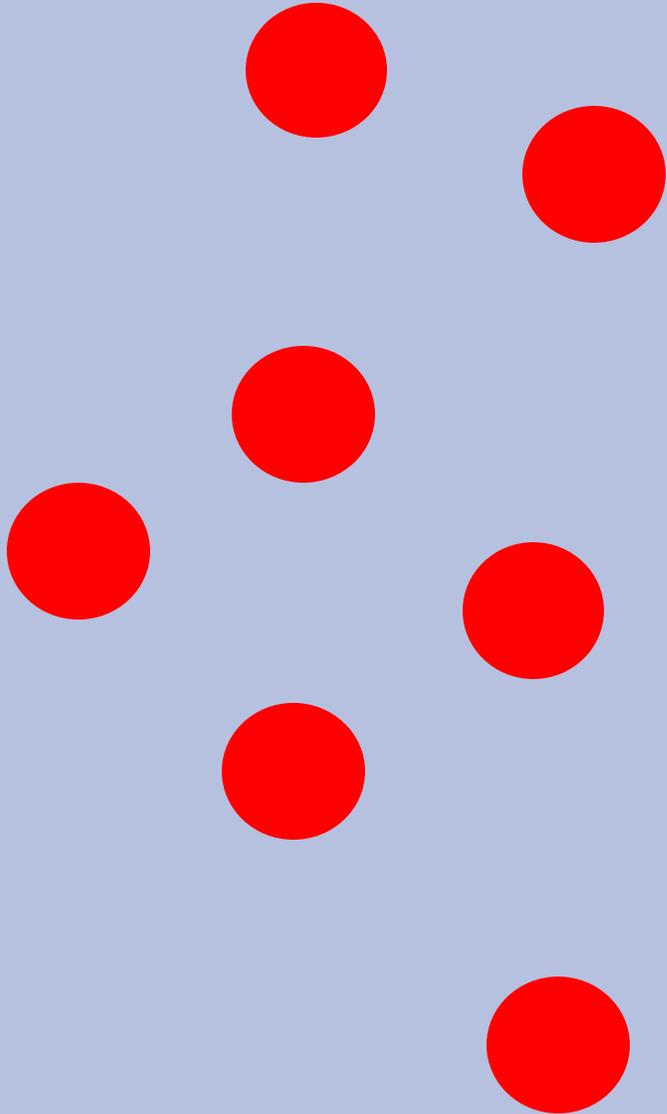
We are learning about

Mixtures

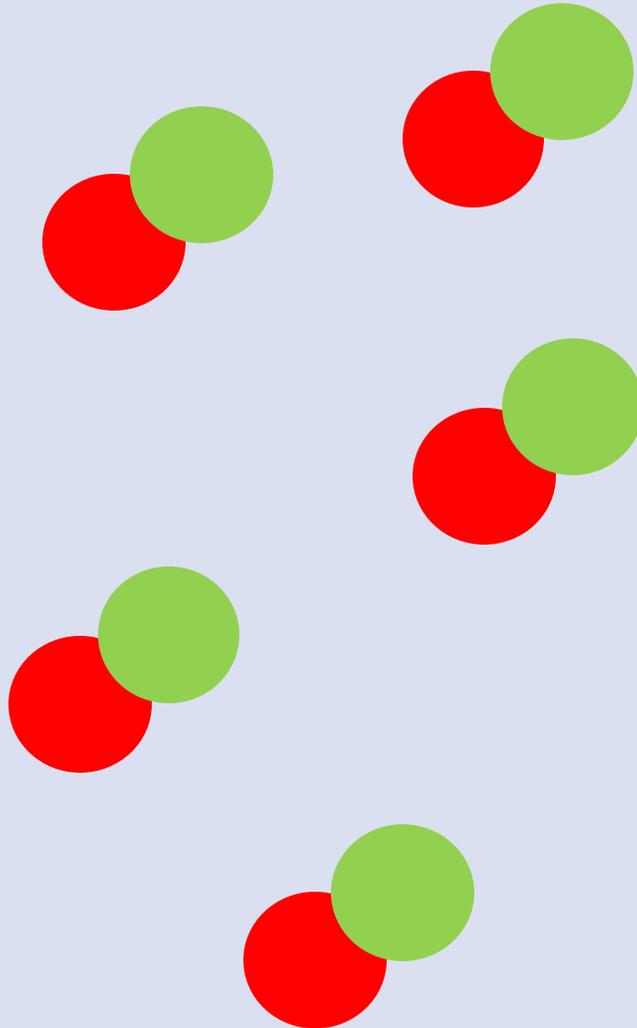


- I can describe the words mixture
- Give examples of every common mixtures.
- Describe how the technique of using a magnet would separate mixtures.

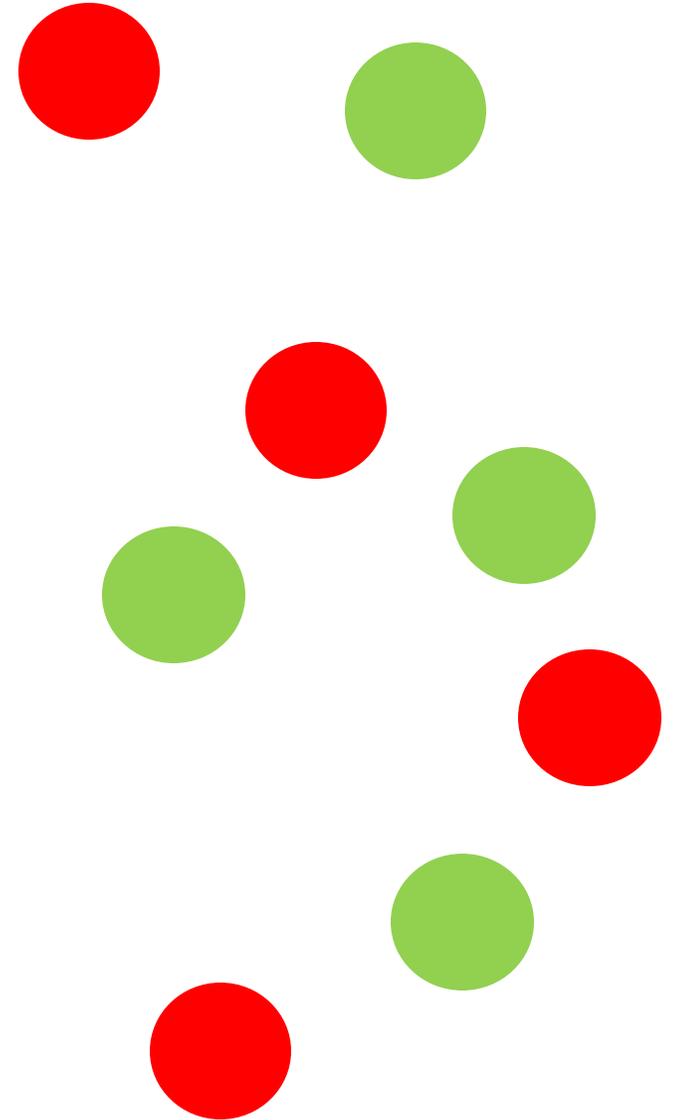
Element



Compound



Mixture



A mixture is when two or more different substances come together **without** chemical joining.

Since they are not joined the substances can be **easily separated**.

Common Mixtures

25/01/2026

Air	Gases, such as oxygen and nitrogen.
Sea water	water, salts and sand.

To investigate a method of separating a mixture of Sulphur and iron.

1. Place the iron and Sulphur mixture in a dish.
2. Place a magnet close to the mixture.
3. Observe what element is separated.

Daily Review

1. What is an element?
2. Give an example of an element
3. What is a mixture
4. Give an example of a mixture
5. How can we separate a mixture?



We are learning about

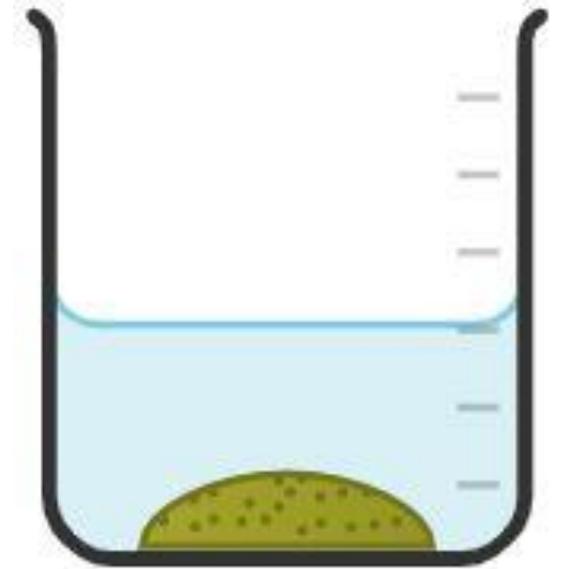
Mixtures



- Describe what technique can be used to separate sand from salty water.
- Describe what can be done in order to separate salt and water.

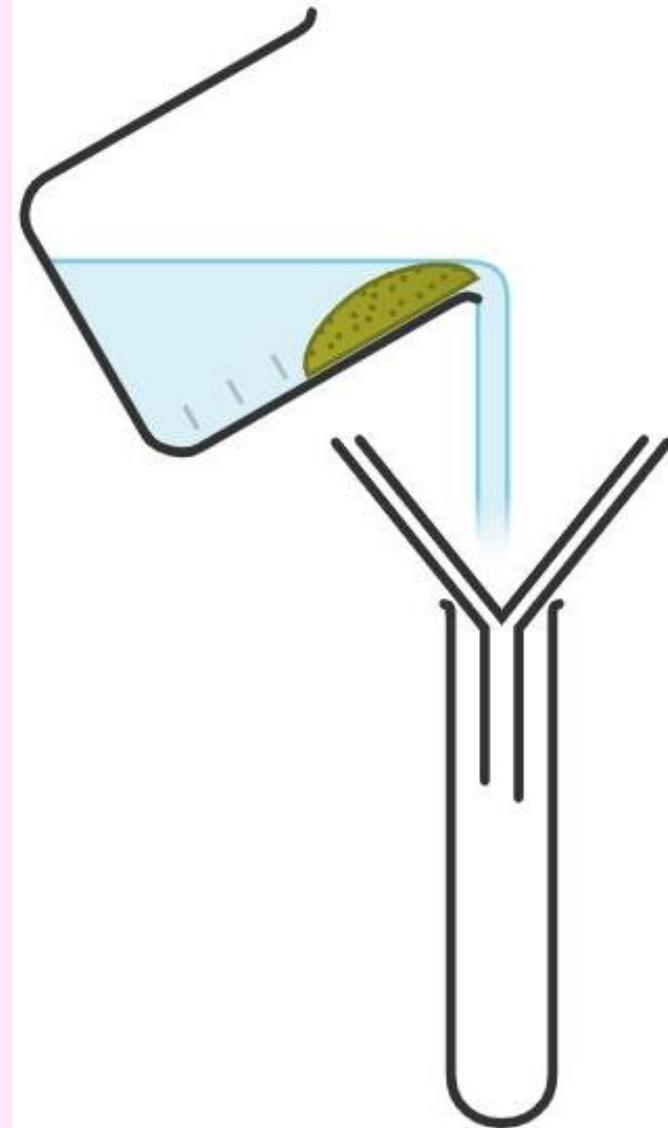
Separating Sand and Salt

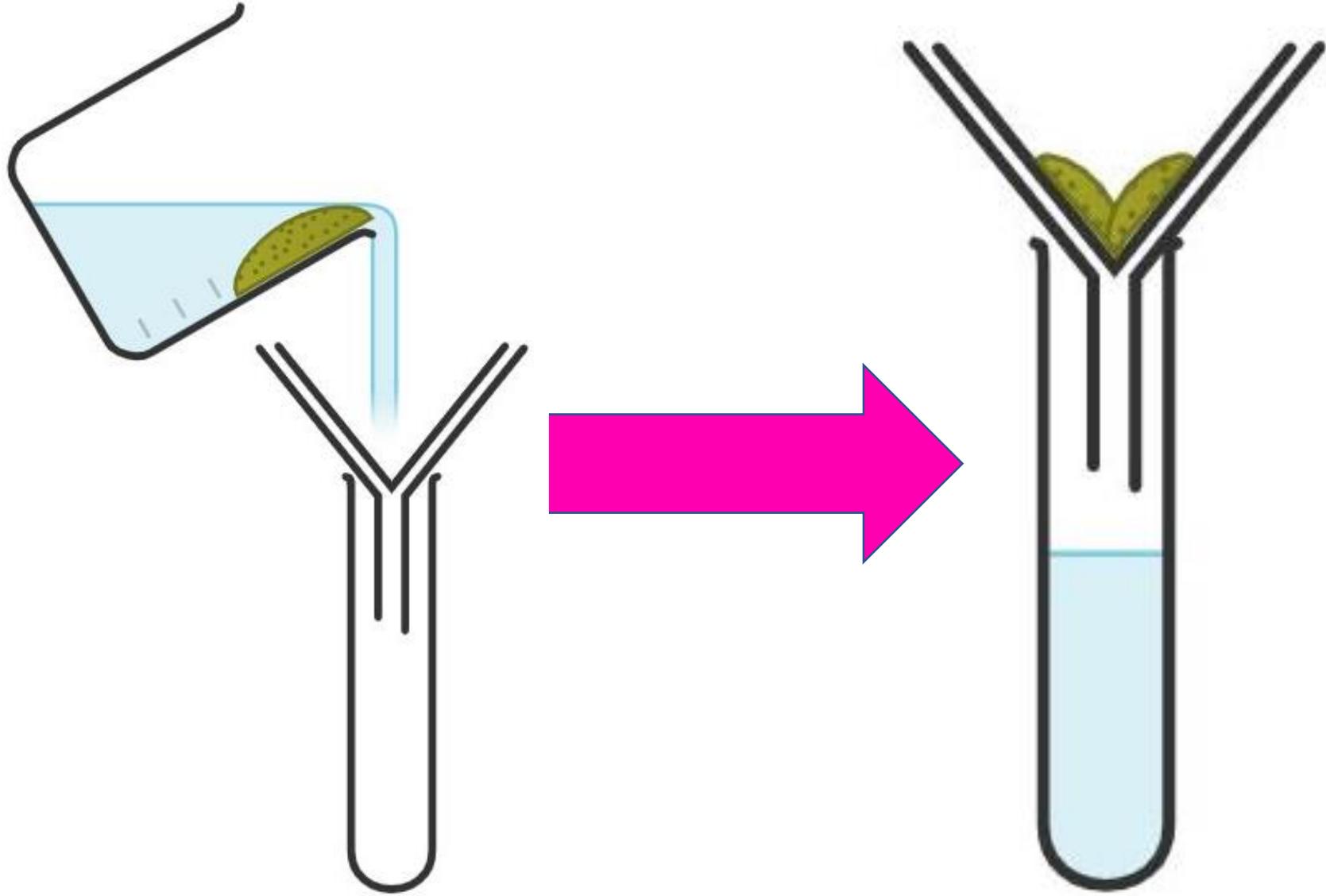
- Fill the beaker with water to the first mark (25cm³)
- Add 4 spatulas of sand/salt mixture and swirl to dissolve the salt.



Separating Sand and Salt

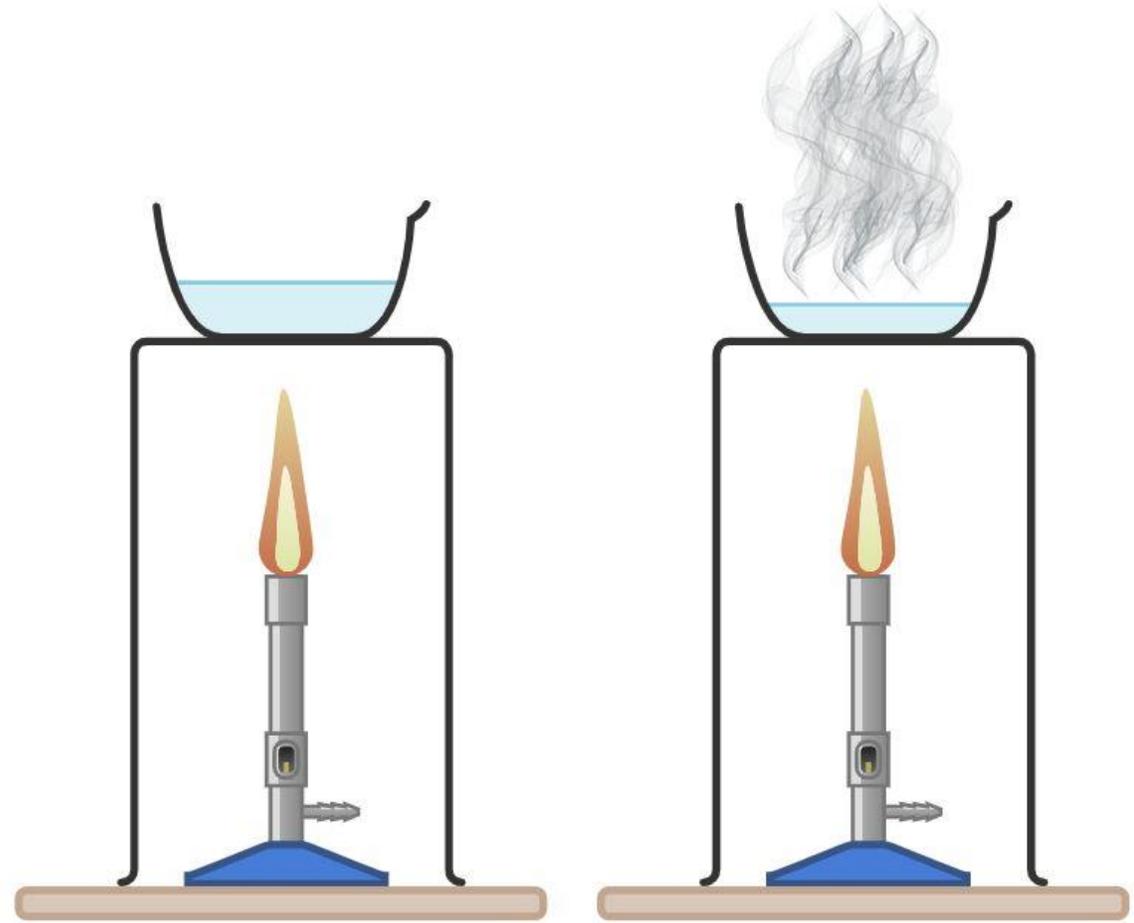
- Put a filter funnel (with paper) into a boiling tube and filter the mixture.
- Set the filter paper with the sand to one side





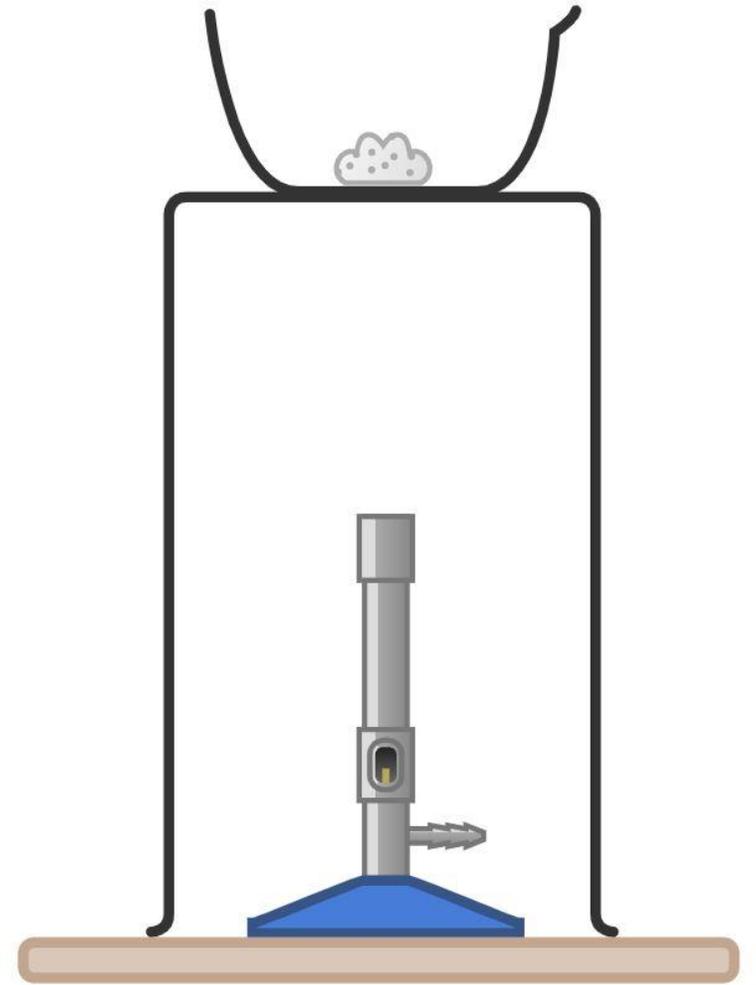
Separating Sand and Salt

- Pour the liquid from the boiling tube into the evaporating basin.
- Heat the liquid to remove the water



Separating Sand and Salt

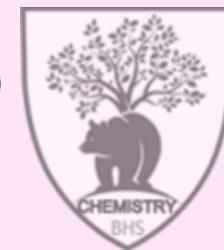
- There should be no water left and only solid salt.
- Salt can be separated from water by boiling and evaporating.



- Magnets
- Filtering
- Boiling / Evaporating

Daily Review

1. What is an element? give an example
2. What is a mixture? give an example
3. How can we separate a mixture? (3 ways)
4. What did Demetri Mendeleev do?
5. Describe what is 'special' about alkali metals.



We are learning about

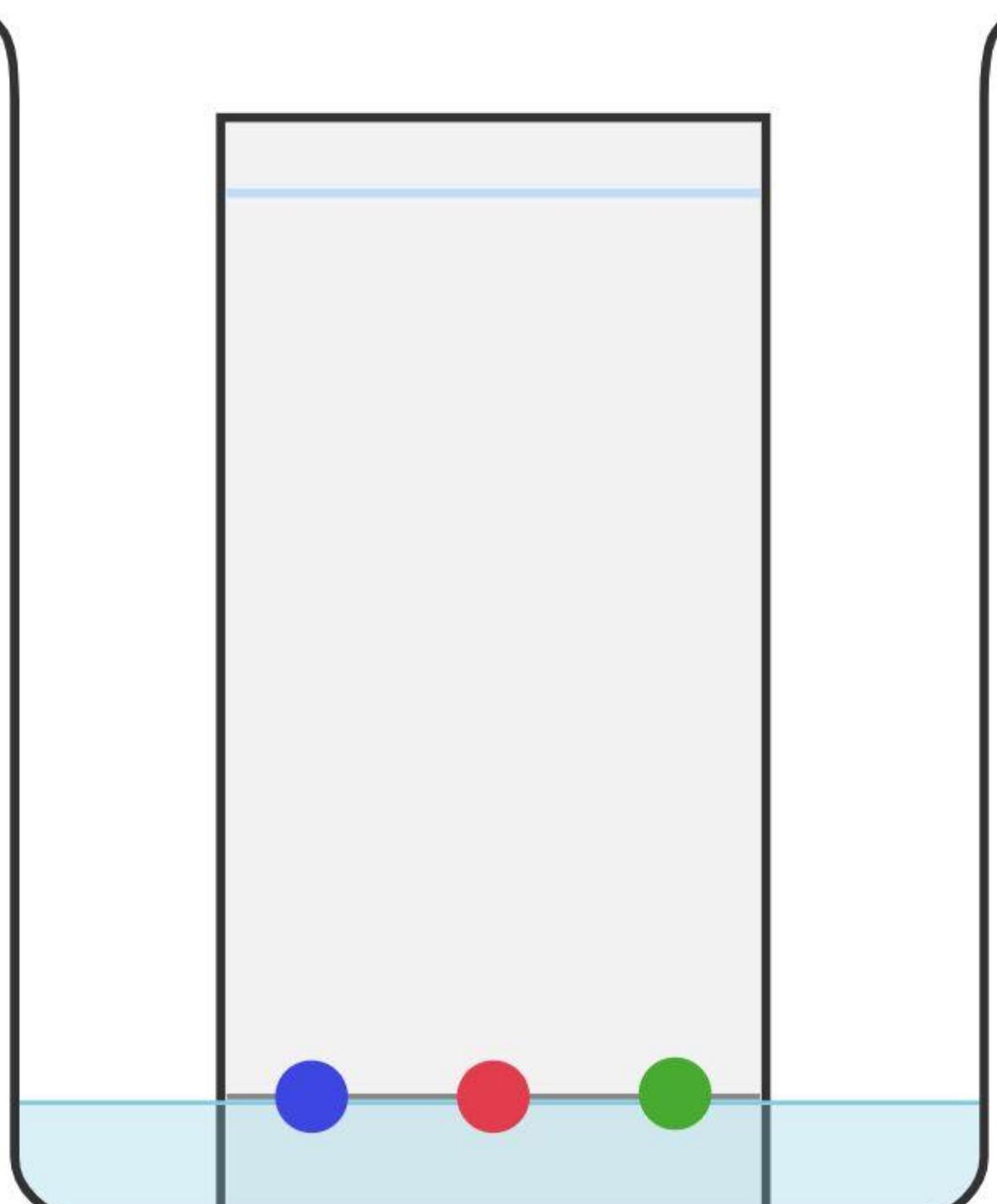
Chromatography



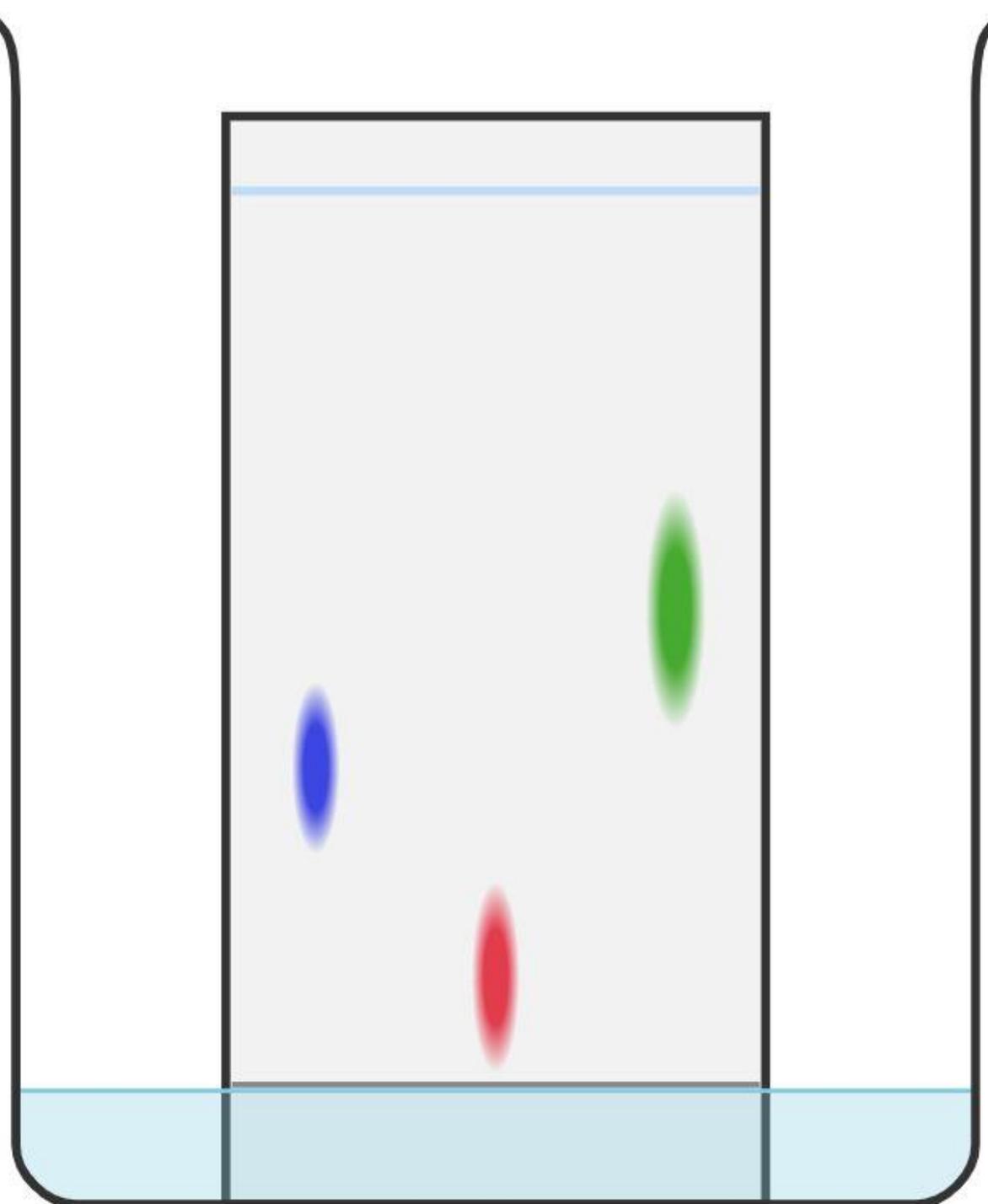
- State what chromatography is used for.
- Describe how chromatography can be used to separate mixtures.

Draw a small pencil line 1cm away from the bottom.

Colour small circles spaced out on the line using different coloured pens.

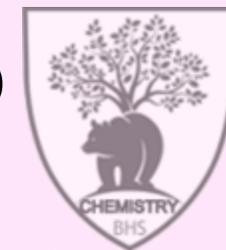


Coloured ink in pens are made from different coloured dyes. The dyes can be separated using chromatography.



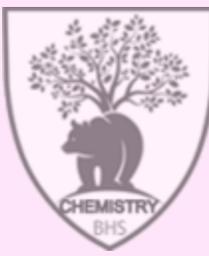
Daily Review

1. What is an element? give an example
2. What is a mixture? give an example
3. How can we separate a mixture? (3 ways)
4. Describe what is 'special' about alkali metals.



We are learning about

Compounds



- Define what is meant by a compound.
- Compare the differences between elements, compounds and mixtures.
- State the names of elements found within named compounds.

Compounds and Mixtures

25/01/2026

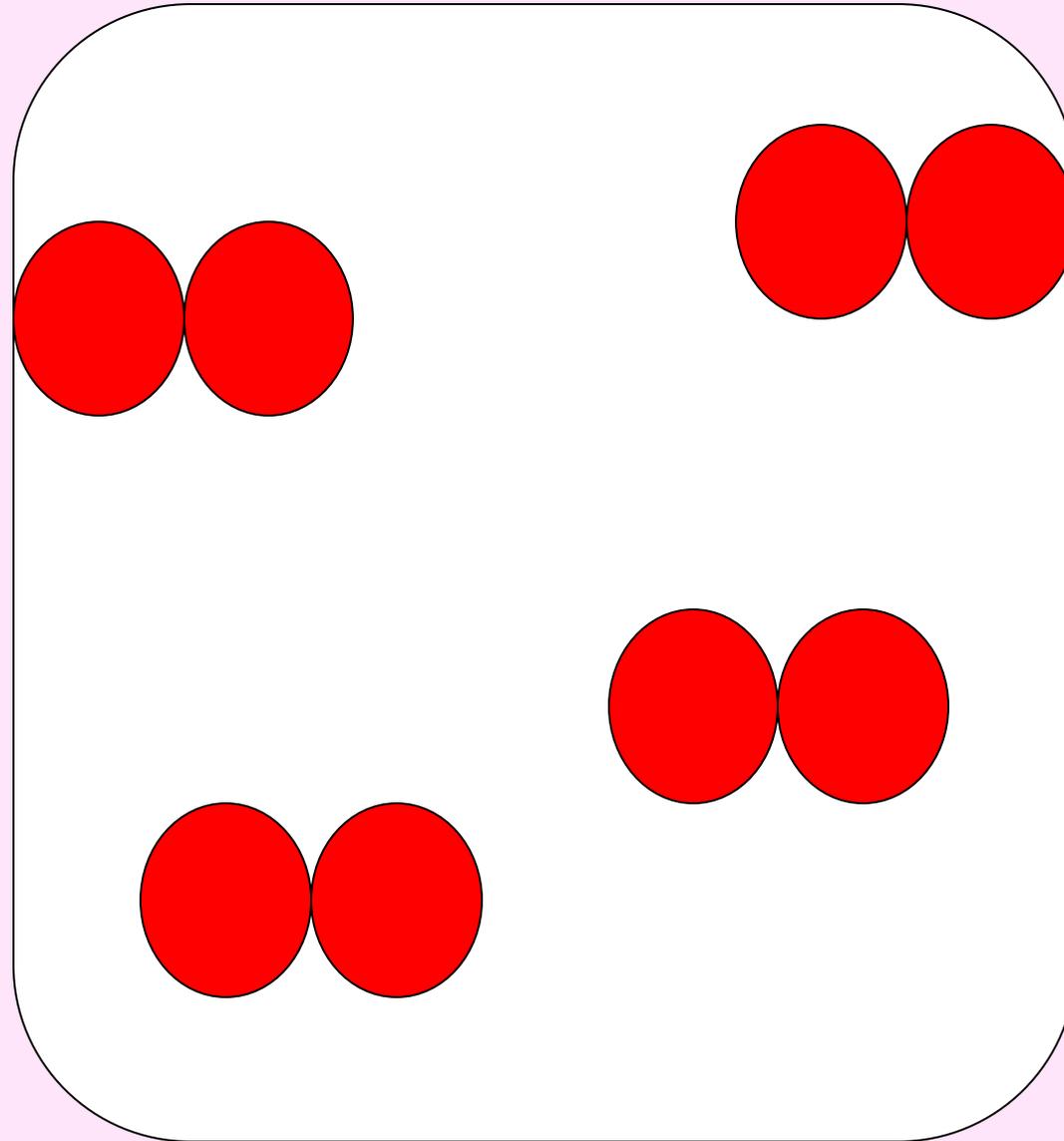
Compound:

2 or more different elements
joined together

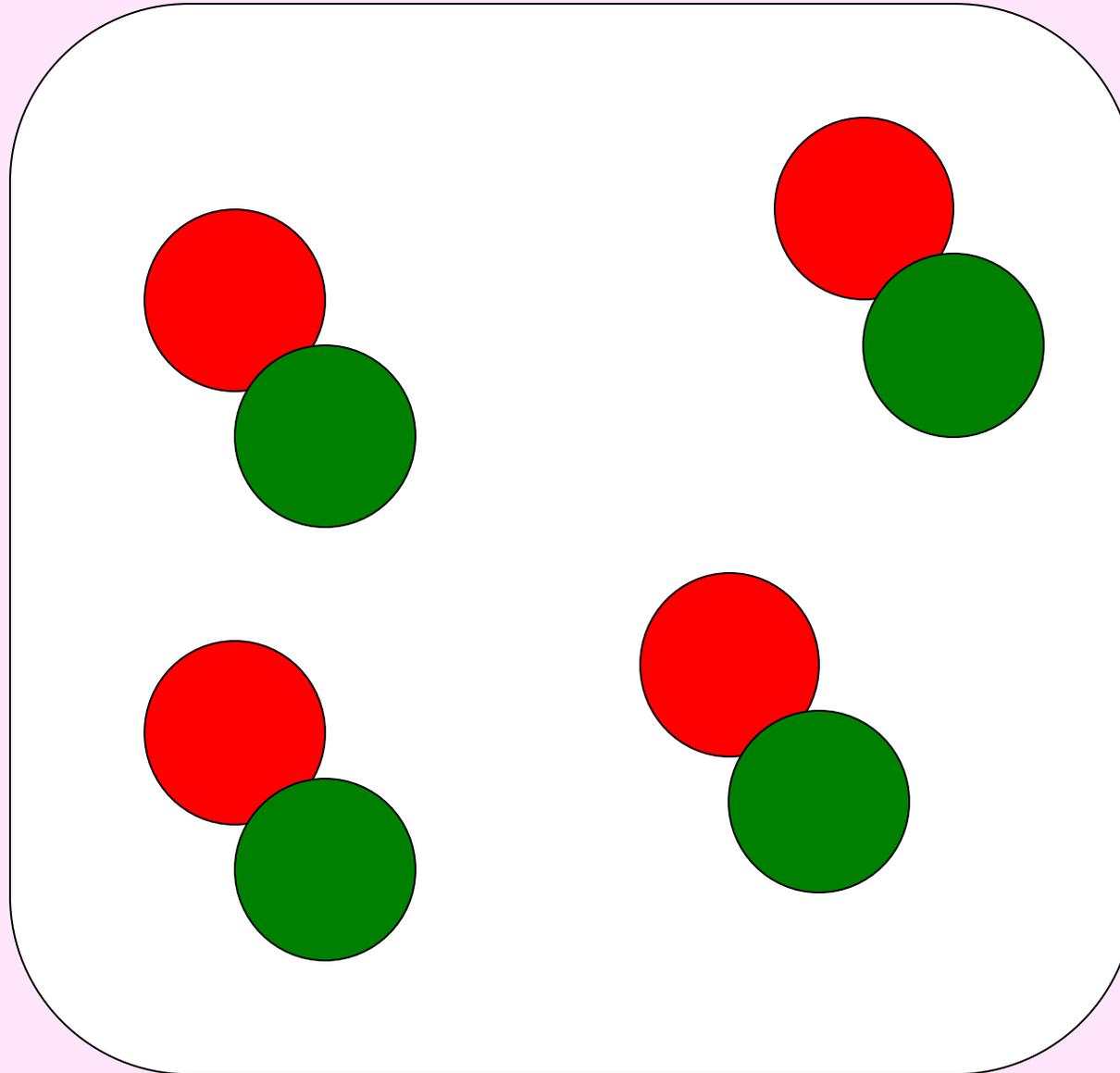
Mixture:

2 or more elements mixed but
not joined

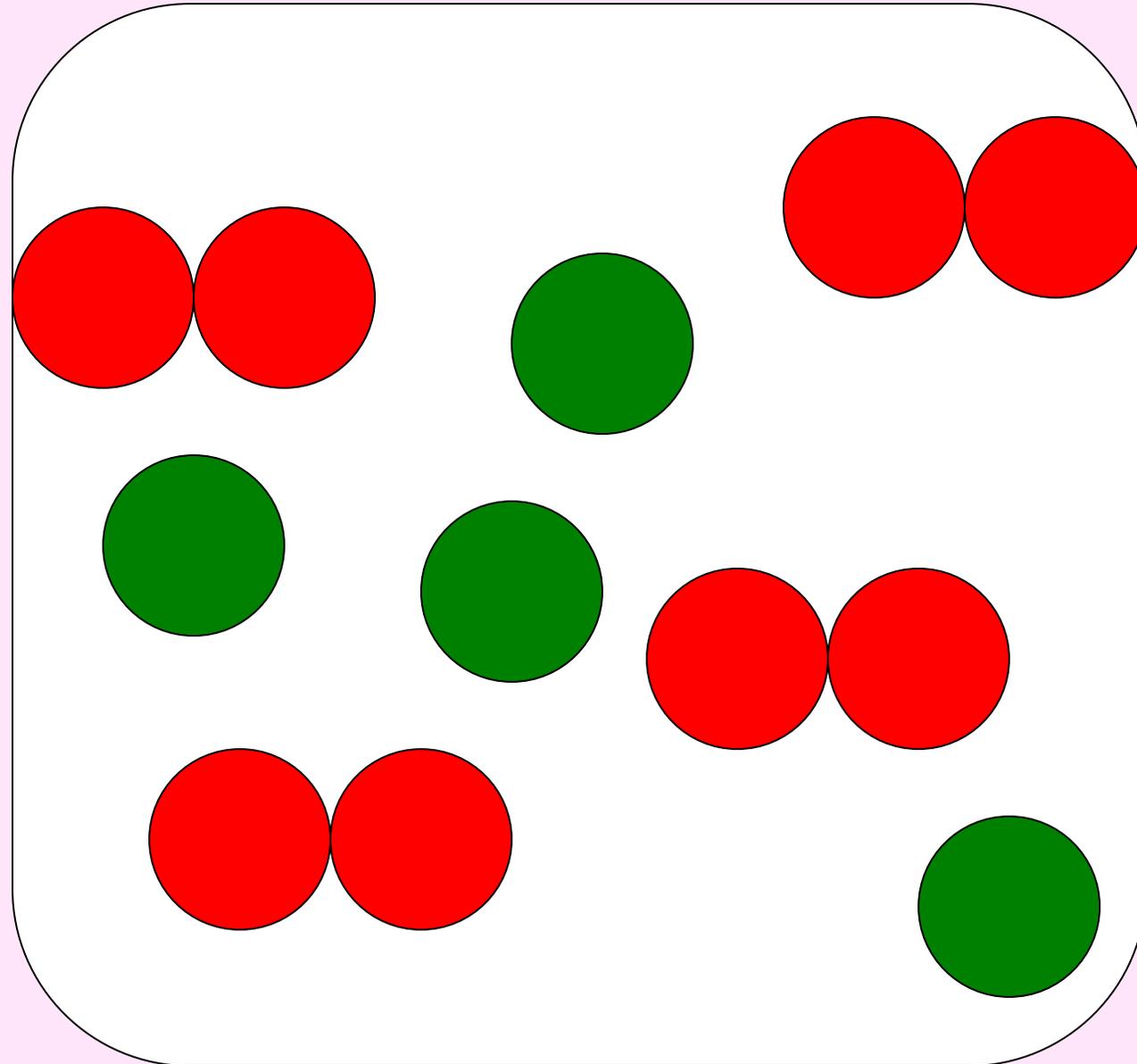
Compound, Element, or Mixture



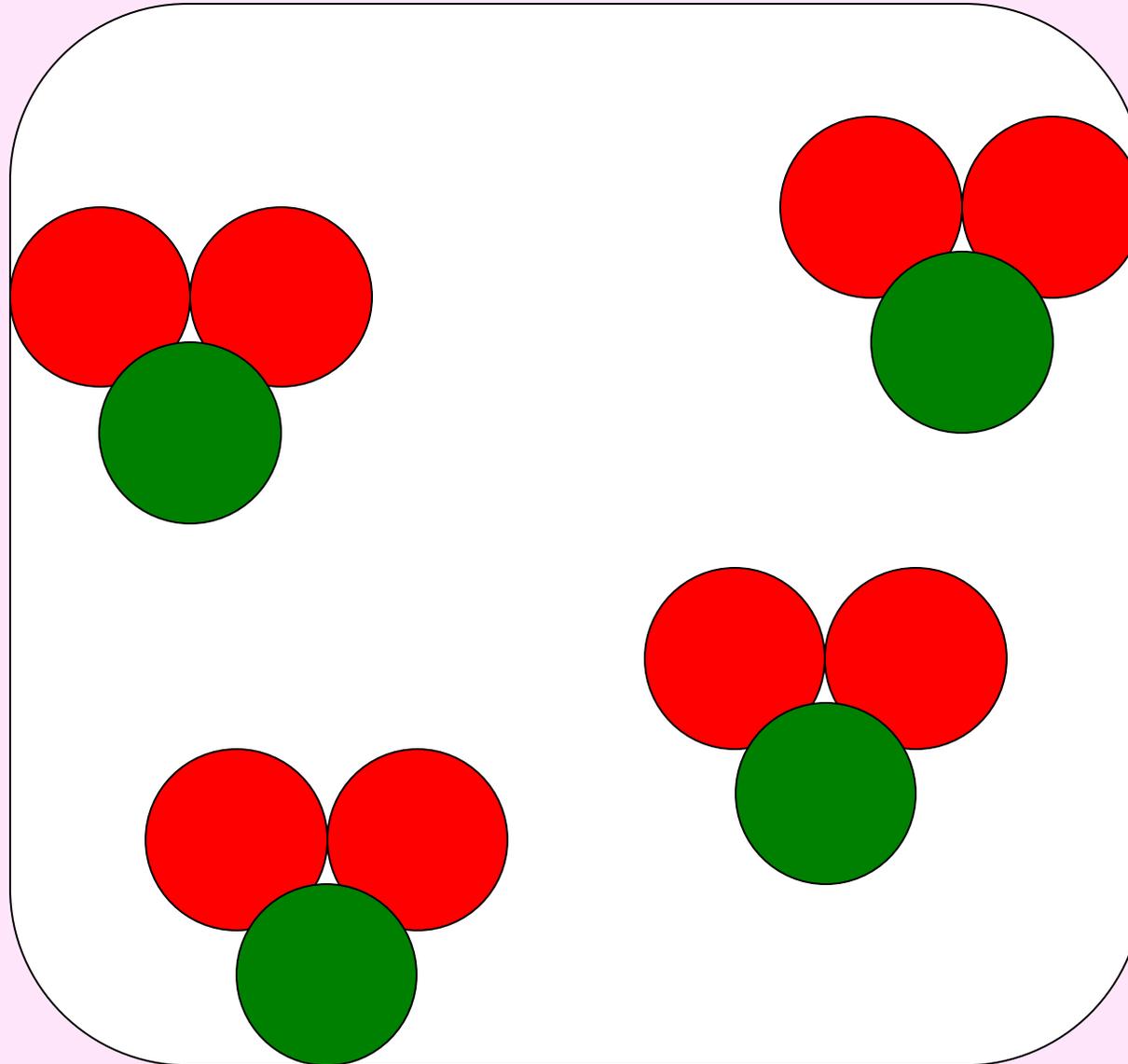
Compound, Element, or Mixture



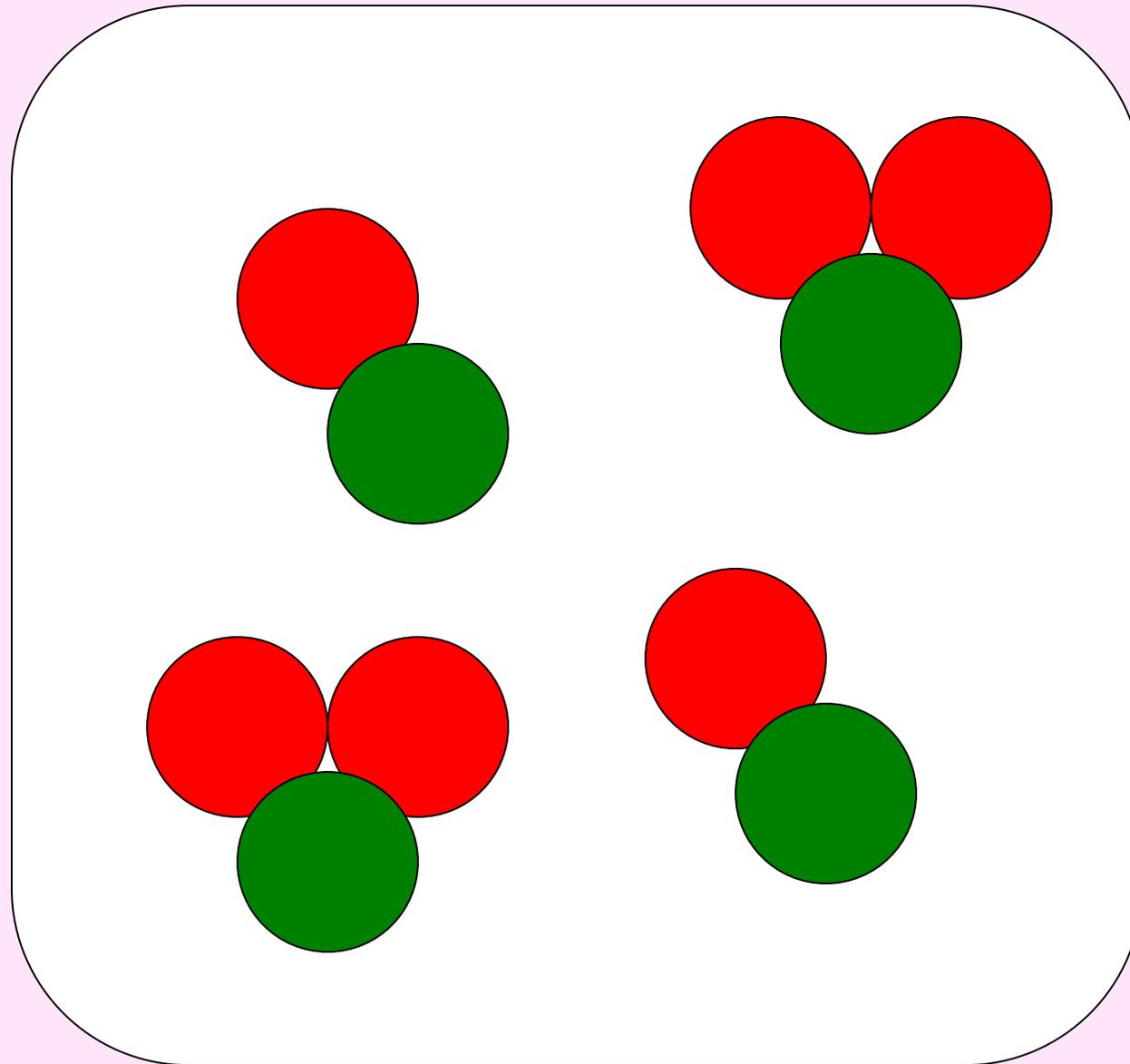
Compound, Element, or Mixture



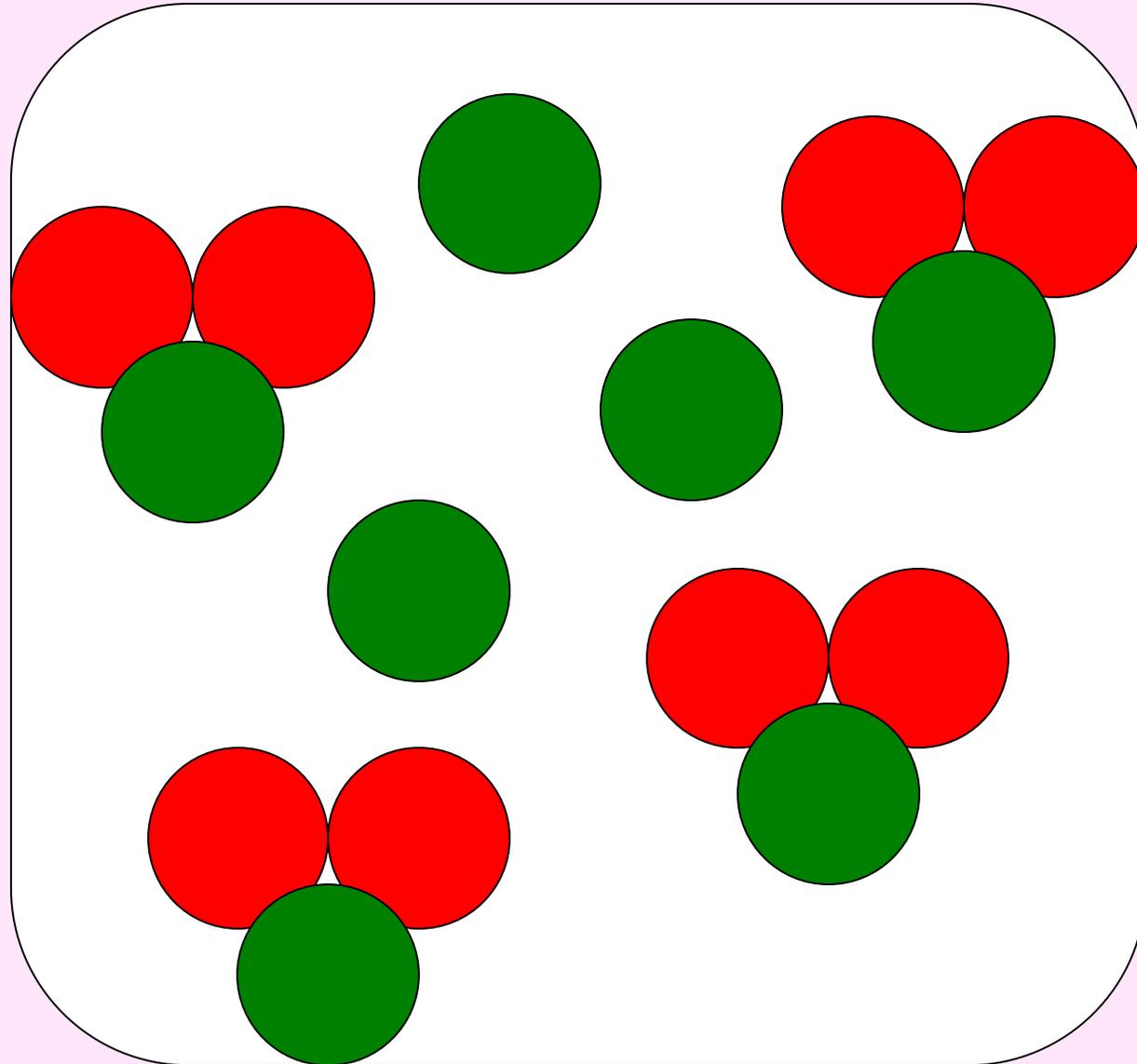
Compound, Element, or Mixture



Compound, Element, or Mixture



Compound, Element, or Mixture





Word Equation:

Iron+Sulphur \rightarrow Iron Sulphide

Examples:

Water, Carbon dioxide, Salt, Sugar

Properties of iron Sulphide are different to iron and Sulphur.

This is true for all Compounds

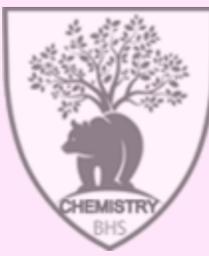
Task

1. QR Code

2. L5 Compounds

OneNote S1 Chemistry,
Periodic Table



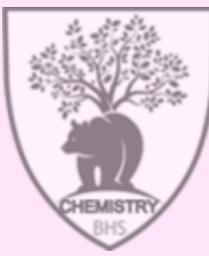


- I can describe the terms compound and mixture
- I can describe how these 3 things are related
- I can identify compounds, elements, and mixtures
- I can give examples of common compounds



We are learning about

Naming Compounds



I can name compounds

I can work out which elements are in a compound

Naming Compounds

25/01/2026

Sodium

and

Chlorine

Element 1

Element 2 ide

Sodium

Chloride

Name the Elements

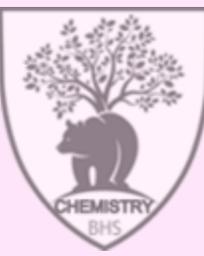
1. Sodium Chloride
2. Magnesium Oxide
3. Lead Iodide
4. Silver Bromide

Task

Practice Test

Revise for topic Test

Textbook (P68-83)



We are

**Revising Periodic Table,
Compounds and Mixtures**